ST. XAVIER'S COLLEGE
(Autonomous), Ahmedabad

Reaccredited 'A+' (CGPA:3.27 out of 4) by NAAC (4th Cycle) | Affiliated to Gujarat University

Syllabus for

Grant in Aid

Master of Science (M.Sc.) – Organic Chemistry

(With modifications upto 15th September, 2022)

Programme Specific Outcome

A student completing this program will be able to

PSO1: Knowledge: Apply the basic knowledge of inorganic, organic, physical and analytical chemistry, in the various fields of pharmaceutical sciences, environmental sciences, forensic sciences and biochemical analysis

PSO2: Laboratory skills: Apply various techniques for the qualitative and quantitative analyses based on instrumental and physico-chemical methods as well as apply the principles of synthetic organic chemistry for organic preparation.

PSO3: Environmental concern: Become aware about various kinds of environmental pollution and apply the principles of green chemistry in industrial and laboratory processes

PSO4: Employability/future prospects: Develop analytical and problem- solving skills as well as apply good laboratory practices necessary for performing various activities in the industry

PSO5: Scientific communication: To develop scientific *communication skills, thereby* the students are able to express ideas clearly and convincingly, in written and oral forms.

Table of Courses

St. Xavier's College (Autonomous)

Ahmedabad-09

M.Sc. Organic Chemistry

Semester: 1
Course Name: Inorganic Chemistry (Theory) **Course Code: PCH-1801**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:

CO1: Apply the approximation methods in quantum theory to solve the Schrodinger wave equation for multielectronic atomic systems such as helium

CO2: Use molecular symmetry and group theory to simplify problems involving molecular properties

CO3: Remember the basic laws and mechanisms of magnetochemistry and recognize different magnetic materials

CO4: Remember the structure and functioning of natural bio-inorganic molecules as well as appraise the role of synthetic coordination compounds in therapy

Unit 1 -Quantum theory and Atomic Structure (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Quantum theory and Atomic Structure]

Postulates of quantum mechanics, setting up of different observables, commutation relations, step-up and step-down operators. Simple harmonic oscillator. Angular momentum of inner quantum number j. Approximation methods: Variation method and application to Hatom. Perturbation theory (first order and non-degenerate), application to the Heliumatom.

Unit 2- Symmetry and Group Theory (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Symmetry and Group Theory]

Representation of groups –some properties of matrices & vectors, representation of groups, the Great orthogonality theorem and its consequences, character table, wave functions as basis for irreducible representations, direct product, identifying non- zero matrix elements.

Unit 3- Magnetochemistry (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Magneto chemistry]

Magnetic susceptibility and basic derivation of diamagnetic susceptibility, pascal constant and its utility, Curie law and Curie-Weiss law, antiferromagnetism and ferromagnetism. Types of ant ferromagnetism, antiferromagnetic exchange pathway: Direct–metal-metal interaction and Indirect-atom exchange i.e. super exchange mechanism.

Unit 4- Bio-inorganic Chemistry (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Bio-inorganic Chemistry]

Metalloporphyrins (enzymes) definition, hemoglobin and myoglobin, cytochrome, vitamin B_{12} (cyanocobalamin), zinc metallo enzymes, nitrogen fixation, essential and trace elements in biological system, biochemistry of non metals K, Na pump (action of bath ions), toxic metals and their toxicity.
Co-ordination compounds in medicine
Chelation therapy, gold compounds and rheumatoid arthritis, anticancer drugs –platinum complexes, gold complexes, metallocenesetc, antimicrobial agents, metal complexes as radio diagnostic agents, magnetic resonance imaging.

Reference Books: PCH 1801: Inorganic Chemistry (Theory)

Core Reference Books:

(1) Introduction to Quantum Chemistry, A. K. Chandra, Tata MacGraw Hill
(2) F. A. Cotton, Chemical Applications of Group theory, Wiley Eastern 2nd Edn.1992
(3) Quantum Chemistry by R. K. Prasad, New Age International Publishers (1985)
(4) Elements of Magnetochemistry, Dutta and Syamal, 1993
(5) Bioinorganic Chemistry, I. Bertini, H. B. Gray and S. J. Lippard

Other Reference Books: PCH 1801: Inorganic Chemistry (Theory)

(1) Quantum Chemistry, Ira N. Levine, Prentice Hall
(2) Elementary Quantum Chemistry by D. L. Pilar, McGraw Hill Book Co, New York (1968)
(3) Quantum Mechanics in Chemistry, M. W. Hanna The Benjamin Pub.
(4) Molecular Quantum Mechanics, Third Edition, P. W. Atkins and R.S. Friedman
(5) Group theory and symmetry in chemistry, L. H. Hall (McGraw Hill)
(6) Group theory in Chemistry V. Ramkrishnan & M. S. Gopinadhan Vishal Pub.1996.
(7) Inorganic Chemistry, Alan G. Sharpe Third Edition,
(8) Theoretical Inorganic Chemistry, M. C. Day, J. Shellin
(9) Hermann Dugas, Bioorganic Chemistry, A Chemical Approach to Enzyme Action, Springer International Edition
(10) Inorganic Chemistry, K. F. Purcell and J. C. Kotz.
(11) Principles of Bioinorganic Chemistry, S. J. Lippard and J. M. Bers
(12) Bioinorganic Chemistry, I. Bertini, H. B. Gray, S. J. Lippard and J. S. Valentine, University Science Books
(13) Introduction to Magnetochemistry, Alan Earnshaw, 1968

St. Xavier's College (Autonomous)
Ahmedabad-09
M.Sc. Organic Chemistry

Semester: 1
Course Name: Organic Chemistry (Theory) **Course Code: PCH-1802**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Apply the knowledge of basic organic reaction mechanisms to infer various aspects of elimination reaction, nucleophilic reactions and molecular rearrangements

CO2: Remember the generation and reactivity of reactive intermediates like carbocations, carbanions etc

CO3: apply the theories of aromaticity to examine molecules and predict their aromatic and antiaromatic character

CO4: employ the fundamentals of stereochemistry to predict the configuration of complex chiral molecules and their roles in enantioselective synthesis

Unit-1 **(15L)[14Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Elimination Reaction and Nucleophilic Substitution Reaction]

(A) Elimination Reaction

The E1, E2 and E1CB reaction mechanism. reactivity-effects of substrate structures, attacking base, leaving group and medium. Regiochemistry of E1 and E2 elimination reactions. Stereochemistry of E2 eliminations in cycloalkane and related systems. Mechanism and orientation in Pyrolytic eliminations in (1) Acyclic systems (ii) Alicyclic systems (iii) Cope eliminations. Thermal decomposition without rearrangement. - The chugaev reaction.

(B) Nucleophilic Substitution Reaction

Mixed S_N^1, S_N^2 and SET mechanism. Nucleophilic substitution at (i) Allylic carbon (Allylic rearrangements) (ii) An Aliphatic trigonal carbon (the tetrahedral mechanism) and at (iii) A Vinyl carbon. Participation of Neighboring groups in Nucleophilic substitution by (a) Carboxylate anion (b) Halogen atoms (c) Hydroxyl groups (d) Acetoxyl group (e) Phenyl group (f) RS group (g) Participation by π-bond.

Unit-2 Molecular rearrangements **(15L)[14Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Molecular rearrangements]

Introduction: Definition and classification.

(I) Rearrangements Induced by Cationic or Electron Deficient Sites

Molecular rearrangements involving electron deficient carbon:

(i) Wagner- Meerwein
(ii) Pinacol- Pinacolone rearrangement

(iii) Tiffeneau-Demjanov Rearrangement
Rearrangements to Electron Deficient Heteroatoms
Electron deficient Nitrogen:
(i) Lossen rearrangement
(ii) Curtius rearrangement
(iii) Schmidt rearrangement
Electron deficient Oxygen:
(i) Baeyer-Villager Rearrangement

(II) Rearrangements Induced by Bases or Electron Rich Sites

(i) The Favorskii Rearrangement
(ii) Wittig and Stevens Rearrangement
(iii) Benzylic acid rearrangement
(iv) The Sommelet-Hauser rearrangement

Unit – 3 (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to reactive intermediates and aromaticity]

(A) Reactive intermediates

Carbocations (classical and non classical), Carbanion, Carbines, Free radicals and Nitrenes: their stability, structure, generation and fate.

(B) Aromaticity

Aromaticity, aromatic character, Frost circle diagram for cyclobutadiene, benzene and others. Resonance and chemical stabilization-aromatic character based on NMR criteria, Huckels molecular orbital (HMO) method, MO of simple organic systems such as ethene, allyl and butadiene. Aromaticity in benzenoid and non-benzenoid compounds and charged rings, annulenes, fulvenes, azulenes, antiaromaticity and homoaromaticity.

Unit –4 Stereo Chemistry (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Stereo Chemistry]

Planar and helical Chirality: Configurational nomenclature: Planar chiral ANSA compounds and trans-cyclooctene. Helical chiral compounds. Enantioselective Synthesis – Epoxidation of allyl alcohols (Sharplessepoxidation), Enantioselectivity through Hydroboration-Oxidation, Enantioselectivity through use of Phase transfer catalysts, Reduction of ketones with chiral hydride donors. Asymmetric resolution: Dynamic resolution, Dynamic kinetic resolution and Dynamic Thermodynamic resolution.

Reference Books: PCH 1802: Organic Chemistry: (Theory)

Core Reference Books

(1) Advanced Organic Chemistry, Reactions Mechanisms and Structure, J. March, 6thEdition, John Wiley.
(2) Advanced Organic Chemistry Part A: Structure and Mechanism and Part B:Reaction and synthesis ,Francis A. Carey, Richard J. Sundberg, 5th Edition, Springer .
(3) Advanced Organic Chemistry Part B: Structure and Mechanism and Part B:Reaction and synthesis ,Francis A. Carey, Richard J. Sundberg, 5th Edition, Springer
(4) Stereo Chemistry, P.S. Kalsi , New Age Publications.
(5) Organic Reaction mechanism, Third edition, V. K. Ahluwalia, RakeshkumarParashar, Narosa Publishing house New Delhi.

IV. Other Reference Books: PCH 1802: Organic Chemistry:(Theory)

(1) Carbenes, nitrenes and arynes, T.L. Gilchrist and C.W. Rees.
(2) Guidebook to Mechanism in Organic Chemistry by Peter Sykes, 6th Edition, Prentice Hall
(3) Organic Chemistry, Jonathan Clayden, Nick Geeves, Stuart Warren, 1st Edition, Oxford University Press.
(4) Principles of Organic Synthesis, R.O.C. Norman and J.M. Coxon, 3rd Edition, Blackie Academic and Professional.
(5) Reagents in Organic Synthesis- Fieser and Fieser, John Wiley.
(6) Stereochemistry of Organic Compounds, Ernest L. Eliel, Samuel H. Wilen, Wiley-Blackwell.
(7) Organic Chemistry, T.W. Graham Solomons and Graig B. Frymes, John Wiley and Sons.
(8) Dynamic Stereochemistry of Chiral Compounds: Principles and Applications, Christian Wolf, RSC publishing.
(9) Organic Chemistry Vol 1-2 I.L.Finar 5^{th} edition, ELBS.

St. Xavier's College (Autonomous)
Ahmedabad-09
M.Sc. Organic Chemistry

Semester: 1
Course Name: Physical Chemistry (Theory) **Course Code: PCH-1803**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Employ the fundamentals of thermodynamics to examine the third law as well as interpret the changes from ideal to real states for solutions and gases

CO2: use the principles of chemical kinetics to derive the rate equations for complex and fast reactions as well as remember the experimental techniques used to study the kinetics of these reactions

CO3: employ the fundamentals of solid-state chemistry to examine the bonding in solids and interpret phenomenon such as diffusion and electrical conduction and super conduction in solids

CO4: Use the BET adsorption isotherms to calculate surface area, surface tension etc; to understand heat of adsorption and ways of determining it and using this to understand more about micelles

Unit -1 Chemical thermodynamics (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Chemical thermodynamics]

Experimental verification of third law of thermodynamics, entropy correction for real gases, partial molar quantities and their determination, Gibbs-Duhem equation, chemical potential, chemical potential of idea gases and solutions, Raoult's law, real solutions, free energy and solutions, activity and activity coefficients, fugacity of gases and liquids and methods of its determination. Non equilibrium thermodynamics-basic concepts.

Unit -2 Chemical Kinetics (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Chemical Kinetics]

Unimolecular reactions, chain reactions and branched chain reactions, explosion limits, chain reaction between hydrogen and bromine, theory of absolute reaction rates, Kinetics of fast reaction and some experimental techniques for studying fast reactions like NMR, Mass Spectoscopy, Gas chromatography, Flow method and Flash photolysis method.

Unit -3 Solid state chemistry (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Solid state chemistry]

Bondingin solids and electronic structure in solids, bond theory-metals, semiconductors and insulators, defects in crystals, calculation of schottly and Frenkel defects using statistical method,

non-stoichiometry –FeO (wustite), solid electrolytes, diffusion in solids- Fick's laws, mechanism of diffusion, electrical conductivity in solids, super conductivity, perovskites.

Unit -4 Surface chemistry (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Surface chemistry]

Physical and chemical adsorption, BET equation, heat of adsorptionand its measurement by Calorimetric and Clausius Clapeyron equation methods, determination of surface area of adsorbents by BET method , surface tension and adsorption from solutions, Gibb's adsorption equation, micellisation and critical micellar concentration (cmc).

Reference books: PCH1803: Physical Chemistry(Theory)

Core Reference books:

(1) Advanced physical chemistry by Gurdeep Raj 35th revised edition, Goel publishing house.
(2) Thermodynamics for chemists by S.Glasstone, Read Books, 2007.
(3) Solid state chemistry ,An introduction : 4th edition CRC press by Smart and Moore

IVReference books: PCH 1803: Physical Chemistry (Theory)

(1) Physical chemistry by W.J.Moore, 5th edition, orient longman private ltd.
(2) Textbook of physical chemistry by S. Glasstone, D. Van Nostrand company, inc., 1946.
(3) Textbook of physical chemistry by Peter AtkinsJulio de and Paula, 9th edition, oxford press.
(4) Advanced physical chemistry by J.N.Gurtu, A.Gurtu, 11th edition , Pragati prakashan.
(5) Physical chemistry by S. Castellan, 3rd edition, Pearson Custom Publishing.
(6) Thermodynamics of non equilibrium processes- Karapitianeh
(7) Chemical Kinetics by Laidler, 3rd edition, Pearson Education India
(8) Chemical Kinetics – Frost and Pearson
(9) Principles of the Solid State by H.V.Keer, 2nd edition ,New Age Internation (P) Ltd.
(10) Introduction to Solids by L.Azaroff, 1st edition, McGraw Hill Education India Pvt Ltd.
(11) Physical Chemistry of Surfaces by A.W.Adamson,6th edtition, Wiley-Interscience.
(12) Surface chemistry – Osipov
(13) Solid State Chemistry and its Applications by Anthony R West, 2nd edtiton 2014,Wiley.
(14) Chemical Thermodynamics: Classical, Statistical and Irreversible by S Chand, 2nd edition. Rajaram & Kuriakose,

Semester: 1
Course Name: Analytical Chemistry (Theory) **Course Code: PCH-1804**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: identify the sampling technique and calibration method for a given sample/glassware etc. and to learn about validation of analytical methods using different statistical parameters

CO2: Identify, recognize and remember the fundamentals of UV-Vis spectrophotometry and apply them effectively for structural and quantitative analysis of molecules and complexes

CO3: Remember the principle of thermal analysis and recognize the different techniques/instrumentation like DSC, DTA etc for industrial applications

Unit -1Analytical objectives, sampling and calibration methods (15L)[14 Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Analytical objectives, sampling and calibration methods]

Scope of analytical science and its literature, sampling and sample preparation, general steps in chemical analysis, calibration and classification of glassware, validation of analytical methods, finding the best straight line-least square regression ,correlation coefficient, calibration curves, standard addition technique internal standard method. Numericals based on chemical concentrations.

Unit-2 Fundamentals of spectrophotometry (15L)[14 Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Fundamentals of spectrophotometry]

Properties of light, absorption of light, interaction of light with matter and origin of spectra, spectrophotometer-instrumentation of single and double beam, Beers Law-its use, limitation and numericals, photometric accuracy.

Unit-3 Applications of spectrophotometry (15L)[14 Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Applications of spectrophotometry]
Analysis of mixture, measurement of equilibrium constant, Scatchard Plot, Stoichiometry determination-method of continuous variation-Jobbs Plot , Photometric titrations.

Unit-4 Thermal methods of analysis (15L)[14 Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Thermal methods of analysis]
Principle, instrumentation and applications of Thermo Gravimetric Analysis (TGA), Differential thermal analysis (DTA) and Differential Scanning Calorimetry (DSC).

Reference books: PCH1804: Analytical Chemistry (Theory)

Core Reference books:

(1) Principles of Instrumental Analysis, by Douglas A. Skoog, 3rd Edition, Holt- Saunders International Edition.
(2) Quantitative Chemical Analysis, by Daniel C. Harris, 5th Edition, W.H. Freeman and Company, New York.
(3) Fundamentals of Analytical Chemistry by Crouch, West and Skoog,9th edition , Brooks/Cole (2013)

Other Reference books: PCH1804: Analytical Chemistry (Theory)

(1) Analytical Chemistry, by Gary D. Christian, 6th Edition, John Wiley and Sons Inc. New Jersey.
(2) Instrumental Methods of Chemical Analysis, by Galen W. Ewing, 4thEdition, International Student Edition.

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Semester: 1
Course Name: Inorganic Chemistry and Organic Chemistry (Practicals) **Course Code: PCH-1805L**
No. of Credits: 04 **Learning Hours: 2X3 Hours**

Course Outcomes:
CO1: Use the fundamentals of semi-micro qualitative analysis to determine six radicals in an unknown mixture

CO2: Perform green preparation of complexes

CO3: Use the fundamentals of organic synthesis to prepare a variety of organic molecules

CO4: Perform organic estimation of some functional groups

(A) Advanced Inorganic Chemistry (Practicals)
Course Content
(1) Solid phase synthesis of trans-bis glycinato copper (II)
(2) Non-metal complex: Synthesis and characterization of bispyridine iodide nitrate.
(3) Nano-chemistry: Preparation of manganese dioxide nanoparticles.
(4) Synthesis of hexaammine cobalt (III) chloride.
(5) Determine the presence of F, As, Zn, Cd, Pb, Cu in drinking water and heavy metal in food samples.
(6) Determination of the half wave-potential for Cd (II) or Cu (II) or Zn (II) ion in 0.1 M KCl solution.
(7) Study the kinetics of dissociation of tris-O-phenanthroline Fe (II), Ni (II) complex by spectrophotometric method.
(8) Catalytic reduction activity of silver nano particle for p-nitrophenol and other derivatives.

Projects:
(1) Just like heavy metal detection in waste water, in chocolate and toys with the help of nano-particles.
(2) Bio chemical sensing with nano-particles.

References books: PCH1805L (A): Advanced Inorganic Chemistry (Practicals)

(1) Vogel's Qualitative Inorganic Analysis, Revised by G Svehla, Sixth Edition, Longman, 1987.
(2) Monograph on Green Chemistry Laboratory Experiments, Green Chemistry Task Force Committee, DST.

(B) Organic Chemistry (Practicals)

II. Course Content

(a)Preparation of organic compounds: Single Stage Preparations:

(1) Preparation of 1-Phenyl-3-methyl-5 pyrazolone from acetoacetic ester.

(2) Preparation of Dibenzylidene acetone from Benzaldehyde.

(3) Preparation of o-Chlorobenzoic acid from o-amino benzoic acid.

(4) Preparation of 2,4-Dinitroanisole from Anisole

(5) Preparation of Phthalimide from phthalic acid.

(6) Preparation of Para Red

(7) Preparation of methyl Orange.

(8) Preparation of Benzo triazole from o-Phenylene diamine

(9) Preparation of 1,2,3,4 tetrahydrocarbazole from phenyl hydrazine. (Fisher Indole synthesis)

(10) Preparation of p-Bromo acetanilide from acetanilide (Green route)

(b) Quantitative Estimations:

(1) Estimation of ester + acid

(2) Estimation of Formaldehyde

(3) Estimation of glycine

(4) Estimation of amide + acid

References books: PCH1805L (B): Organic Chemistry (Practicals)

(1) A text book of practical organic chemistry – A. I. Vogel
(2) Practical organic Chemistry – Mann and Saunders
(3) A handbook of quantitative and qualitative analysis – H. T. Clarke
(4) Comprehensive Practical Organic Chemistry : Qualitative Analysis V K Ahluwalia& S. Dhingra.
(5) Comprehensive Practical Organic Chemistry : Preparations and Quantitative Analysis V K Ahluwalia& R. Aggarwal Universities Press.
(6) An Advance Course in practical Chemistry, A K. Nad, B. Mahapatra and A. Ghoshal.

Semester: 2
Course Name: Physical Chemistry and Analytical Chemistry (Practicals) **Course Code: PCH-1806L**
No. of Credits: 04 **Learning Hours: 2X3 Hours**

Course Outcomes:
CO1: Use the conductometer to do acid-base titrations involving mixtures

CO2: Use the potentiometer to determine solubility product and oxidation potentials

CO3: Use the pH meter to do acid-base titrations to determine the assay of a drug and concentration in mixtures

CO4: Perform experiments based on adsorption and kinetics

CO5: Perform experiments based on distribution coefficient

CO6: Calibrate glass wares used in an analytical laboratory

CO7: Perform titrations to determine the assay of drugs and other commercial preparations

CO8: Perform water analysis

(A) Physical Chemistry (Practicals)
II. Course Content
(1). Conductometry

(a) To determine on centration of HCl and NH_4Cl in a given solution conductometrically (requirements: 0.05N HCl, 0.5 N NH_4Cl, 0.5 N NaOH)
(b) Estimate the concentration of H_2SO_4, CH_3COOH and C_USO_4 5 H_2O in a given solution conductometrically. (0.005M all against 0.05N NaOH)

(2) Potentiometry
(a) To construct the calibration curve for quinhydrone electrode and hence the standard oxidation potential of quinhydrone electrode. (0.2N CH_3COONa, 0.2NCH_3COOH)
(b) Solubility product of silver halides.

(3) pH metry
(a) To determine the amount of Aspirin in a given solution.(aspirin,0.1N alcoholic KOH, glass and calomel electrodes)
(b) Titration of mixture of bases (Na_2CO_3& $NaHCO_3$) with standard HCl and find the concentration of bases.

(4) Adsorption and kinetics

(a) To study the rate of acid catalysed ionization of acetone in presence of excess acid and acetone at room temp. (requirement: acetone, iodine, sulfuric acid sodium acetate, 0.01sodium thiosulphate, starch)
(b) To determine the autocatalytic reaction between $KMnO_4$ and oxalic acid.

(Requirements:0.1M $H_2C_2O_4$, 10%KI , 0.01M $Na_2S_2O_3$, 0.02M $KMnO_4$,0.2M $MnSO_4$, 1% starch, 1.0 M H_2SO_4)

(5)Distribution method

(a) Distribution of HAC between H_2O and $CHCl_3$ / CCl_4.
(b) Distribution of I_2 between H_2O and CCl_4.

References books:PCH1806L (A): Physical Chemistry (Practicals)

(1) Advanced Practical Physical ChemistyBy J.B.Yadav, 32nd edtion Krishna publication.
(2) Practicals in physical chemistry by P.S.Sindhu, Macmillan 2005.
(3) Experimental physical chemistry by R.C.Das, B.BeheraTata McGraw-Hill, 1983.
(4) Experimental Physical Chemistry by Athwale , Mathur and Parul, 1st edition reprint 2011,New Age International (P) Ltd.

(B)Analytical Chemistry (Practicals)

II. Course Content

(1) Calibration of glass wares and balance.
(2) Determination of %age purity of given sample of Isoniacid.
(3) Determination of %age of Asprin in the given tablet.
(4) Determination of available chlorine in bleaching powder.
(5) Determination of vitamin C in orange juice/amla.
(6) Determination of acetic acid in vinegar.
(7) Determination of sodium carbonate and sodium bicarbonate in washing soda.
(8) Determination of ascorbic acid in vitamin C tablets.
(9) Determination of %age purity of given sample of Analgin tablet.
(10) Determination of calcium and magnesium in water sample.
(11) Determination of sulphate in water sample.
(12) Determination of chloride in water sample.

References books:PCH1806L (B):Analytical Chemistry (Practicals)

(1) Analytical Chemistry Practice, John H. Kennedy, Saunders College Publishing, Second Edition 1990.
(2) Vogels Textbook of Quantitative Chemical Analysis, 6th Edition, 2002.

St. Xavier's College (Autonomous)
Ahmedabad-09
M.Sc. Organic Chemistry

Semester: 2
Course Name: Inorganic Chemistry (Theory) **Course Code: PCH-2801**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Apply the principles of VSEPR and molecular orbital theory to deduce the properties of conjugated systems, conductors, semi-conductors and insulators

CO2: Use the basics of molecular symmetry and spectroscopy to predict the IR and Raman spectra of molecules

CO3: To recognize the bonding modes in organometallic compounds with various organic ligands

CO4: To remember the reaction mechanisms in coordination compounds

Unit 1- Chemical Bonding **(15L)[14Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to VSEPR theory and Chemical Bonding]

VSEPR, Walsh diagrams(tri atomic molecules), Bent rule and Simple Huckel theory of linear conjugated systems, simple Huckel theory of the cyclic conjugated system and aromaticity, self-consistent filed method, valence state ionization potentials,

Band theory of solids, Fermi level, electrical properties, insulators, semiconductors and superconductors (properties).

Unit 2- Application of symmetry **(15L)[14Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to symmetry and IR and Raman spectra]

Application of symmetry to hybrid orbital, molecular orbitals, hybridization schemes for σorbitals, π bonding and molecular orbital for ABn type of molecules.

Application of symmetry to molecular vibrations, interpretation of IR and Raman spectral data.

Unit 3-Organometallic Compounds **(15L)[14Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Organometallic Compounds]

Organometallic compounds of transition elements, stability of metal carbon bond in complexes. Synthesis, uses and structure of organometallic compounds of π bonding organic ligands, 2-electron ligands, olifinic and acetylinic complexes, compound with 3 electron ligand – allylic complexes, compounds. With 4- electron ligands butadiene complexes, n4 complexes of cyclopentadiene, compounds with 5 electron ligands – cyclopantadionyl, compounds with 6 electron ligands, n6 complexes of benzene and its derivatives. Role of organometallic compounds in catalytic reaction.

Unit 4 – Reaction Mechanism (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Inorganic Reaction Mechanism]

Mechanism of substitution reaction in square planar complexes. Kinetics of substitution reaction of platinum (II) complexes.

Effect of leaving group, effect of charge, steric effect, solvent effect, effect of nucleophile, effect of temperature and other effects.

Oxidation-Reduction reaction, electron transfer, tunnelling effect, Marcus –Hush theory, one and two electron transfer inner sphere and outer sphere, effect of ions on rate, electron transfer through extended bridges, unstable oxidation states, hydrated electron.

Core Reference books:

(1) Mechanism of Inorganic Reactions, F. Basolo and R. G. Persons, Wiley Pub
(2) Electrons and Chemical Bonding by H B Gray
(3) Symmetry and group theory by B S Garg

Other References books: PCH2801: Inorganic Chemistry (Theory)

(1) Lectures on Chemical Bonding and Quantum Chemistry, S. N. Datta, A Prism Book
(2) Group theory and symmetry in chemistry, L. H. Hall(McGraw Hill) Coulson's Valence, R. McWeeny, ELBS
(3) F. A. Cotton, Chemical Applications of Group theory, Wiley Eastern 2nd Edn.1992
(4) V. Ramkrishnan& M. S. Gopinadhan, Group theory in Chemistry Vishal Pub.1996
(5) Inorganic Chemistry, Third Edition, Alan G. Sharpe
(6) Theoretical Inorganic Chemistry, M. C. Day, J. Shellin
(7) Chemistry, Fifth Edition, John E. McMurry, Robert C. Fay
(8) An Introduction to Theoretical Chemistry, Jack Simons, Cambridge
(9) Progress in inorganic Chemistry, Vols 18 and 38 ed. J. J. Lippard, Wiley
(10) Reaction Mechanism of Coordination Compounds, C. H. Langford and H. B. Gray
(11) Inorganic Reaction Mechanisms, M. L. Tobe, Nelson Pub
(12) Inorganic Chemistry, K. F. Purcell and J. C. Kotz.
(13) Principles of Bioinorganic Chemistry, S. J. Lippard and J. M. Bers
(14) Mehrotra R. C. and Singh A. Organo Metallic Chemistry, Willey Eastern Ltd., New Delhi
(15) Coates G. E. Green MIH Wade, K and Aylett B. J. Organo Metallic Comounds Chapman and Hall, London

Semester: 2
Course Name: Organic Chemistry (Theory) **Course Code: PCH-2802**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: To recognize the selectivity and utility of a variety of reagents in organic reactions

CO2: To remember the mechanism and synthetic applications of name reactions

CO3: To apply the fundamental knowledge of heterocycles to elucidate the synthesis, structure and application of polyheteroatomic and benzofused heterocycles

Unit - 1: Application Redox and reagents in organic synthesis. (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Oxidation and reduction]

(A) Oxidizing agents, Reducing agents and Mechanistic explanation of oxidation and reduction with example.
Oxidation with Manganese, $KMnO_4$,Chromium, Peracid, Peroxide, Dimethyl dioxarane SeO_2,NBS,DDQ,Chloranil and Oppenauer oxidation.

(B) Reduction:
(i) Reduction with Hydride transfer reagents like $LiAlH_4$,$NaBH_4$,Diborane.
(ii) Reduction by Dissolving metals Zn, Li, Na.
(iii) Birch reduction and catalytic reduction.

Unit – 2 (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Photochemistry and Heterocyclic compounds]

(A) Photochemistry:

(1) Photochemical reactions: Principles of energy transfer, electronic excitation and molecular orbital view of excitation, excited states and excitation and molecular orbital view of excitation, excited states and fate of excited molecules (modified Jablonski diagram), Photosensitization.

(2) Photochemistry of carbonyl compounds: Representation of excited states of ketones, photo reduction Norrish type I & Π reactions, Reactions of cyclic Ketone, oxetane formation (Paterno-Buchi reaction)

(3) Di-π methane rearrangement, Dienone photochemistry, cis-trans isomerisation and photochemistry of conjugated olefins.

(B) Chemistry of Heterocycles

(1) Nomenclature of heterocycles: Few examplesof systematic nomenclature (Hantzsch-Widman system) for monocyclic, fused and bridged heterocycles.
General chemical behavior of following aromatic heterocycles: their synthesis and important applications. (Three examples each)
(2) Five-membered and benzo fused five member heterocycles :Oxazole, Isoxazole, Thiazole, Pyrazole, Imidazole, Benzothiazole and Benzimidazole.
(3) Six membered and benzofused six membered heterocycles :Pyrazine, Pyridazine, Pyrimidine, Cinnoline, Quinazoline, Quinoxaline, Phenoxaline.

Unit – 3Name reactions: (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Name reactions]

General nature, method, mechanism and synthetic applications of the following named organic reactions:

(i)	Suzuki reaction	(ii)	Buchwald Hartwing reaction (cross coupling)
(iii)	Sonogarshira coupling	(iv)	Vilsmeier-Haack reaction
(iv)	Mitsunobu reaction	(vi)	Stobbe condensation
(vii)	Jones oxidation	(viii)	Swern oxidation reaction
(ix)	Michael addition	(x)	Dickmann reaction
(xi)	Knoevanagel reaction	(xii)	Darzen'sglycidic ester synthesis
(xi)	Mannich reaction	(xiv)	Witting reaction

Unit-4 Reagents in organic synthesis: (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Reagents in organic synthesis]

Mechanism selectivity and utility of following reagents:

(i)	Gilman's reagent-Lithium dimethylcuprate	(ii)	Lithium diisopropylamide (LDA)
(iii)	Dicyclohexylcarbodiimide (DCC)	(iv)	1,3 – Dithiane (Umpolung reagent)
(v)	Dess- Martin periodinane	(vi)	Bakers yeast
(vii)	Diisobutylauminiumhydride (DIBAL –H)		
(viii)	Sodium cyanoborohydride ($NaBH_3(CN)$)	(ix)	Grignard reagents
(x)	Sodium borohydride (xi) DDQ	(xii)	n-Butyl lithium

(xiii) Phase transfer catalysis : Quaternary ammonium and phosphonium salts, crown ethers.

References books: PCH2802: Organic Chemistry (Theory)

Core Reference books:

(1) Organic Chemistry, T.W. Graham Solomons and Graig B. Frymes, John Wiley and Sons
(2) Advance organic chemistry by Jerry March
(3) Photochemistry and Pericyclic Reactions by Jagdamba singh and Jaya singh NEW AGE; 3rd edition (1 January 2012)

(4) Organic Chemistry Vol 1-2 I.L.Finar 5th edition, ELBS.

OtherReferences books: PCH2802: Organic Chemistry (Theory)

(1) Modern Synthetic Reactions, H.O.House, W.A. Benjamin.
(2) Principles of Organic Synthesis, R.O.C. Norman and J.M. Coxon, 3rd Edition, Blackie Academic and Proffessional.
(3) Introductory Photochemistry, A.Cox and T.Camp, McGraw Hill.
(4) Photochemistry, R.P. Kundall and A. Gilbert, Thomson Nelson.
(5) Organic Photochemistry, J. Coxon and B. Halton, 2nd Edition , Cambridge University Press.
(6) Strategic Applications of Named Reactions in Organic Synthesis, Laszlo Kurti and Barbara Czak, 1st Edition, Acedemic Press.
(7) Name Reactions and Reagents in Organic Synthesis, Bradford P. Mundy, Michael G. Ellerd, Frank G. Favaloro, 2nd Edition, Wiley – Interscience.
(8) Name Reactions. A Collection of Detailed Reaction Mechanisms., Jie Jack Li, 3rd Edition Springer.
(9) Heterocyclic Chemistry, volume 1-3, R.R. Gupta, M. Kumar and V. Gupta, Springer-Verlag.
(10) Heterocyclic Chemistry, J.A. Joule, K.Mills, and G.F. Smith, 3rd Edition, Chapman and Hall.
(11) Heterocyclic Chemistry, T.L. Gilchrist, Longman Scientific Technical.
(12) Contemporary Heterocyclic Chemistry, G.R. Nikome and W.W. Poudler, Wiley.
(13) Comprehensive Heterocyclic Chemistry, A.R. Kartizky, and C.W. Rees.
(14) Encyclopedia of Reagents for Organic Synthesis, Leo A. Paquette, David Crich and Phillip L. Fuchs, John Wiley and Sons Inc. .
(17) Guidebook to Mechanism in Organic Chemistry by Peter Sykes, 6th Edition, Prentice Hall.
(18) Advanced Organic Chemistry Part A: Structure and Mechanism and Part B:Reaction and synthesis ,Francis A. Carey, Richard J. Sundberg, 5th Edition, Springer .
(21) Advance organic chemistry by Carey and Sundberg,
(22) Advance organic chemistry by Francis A. Carey.

Semester: 2
Course Name: Physical Chemistry (Theory) **Course Code: PCH-2803**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1:To identify the principles of statistical thermodynamics and hence apply statistical mechanics to molecular problems of chemical thermodynamics

CO2: To identify the various aspects of nuclear chemistry to define its application in energy generation and therapeutics

CO3: use the basic principles of physical chemistry to determine the kinetics, thermodynamics and molecular mass of polymers

CO4: Apply the fundamentals of overvoltage, decomposition potentials and electrochemical polarization to explain the theory and working of polarography

Unit-1 Statistical thermodynamics (15L)[14Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Statistical thermodynamics]

Concepts of distribution of molecules, microstate and macro state. Ensemble averaging, Canonical, grand canonical and microcanonical ensembles, Maxwell-Boltzman distribution laws (using Lagrange's method of undetermined multipliers). Fermi-Dirac statistics – distribution law and Bose-Einstein statistics – distribution law.
Partition functions – translational, rotational, vibrational and electronic partition

Unit-2 Nuclear chemistry (15L)[14Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Nuclear chemistry]
Nuclear properties-nuclear radius, coulombic and nuclear potential radius, nuclear spin and angular momentum, magnetic moment, nuclear binding energy, nuclear models-shell model, liquid drop model, Fermi gas model, collective model, radioactive decay, nuclear reactions, evaporation, spallation, fragmentation, fission and fusion reactions, accelerators, reaction cross section, use of radioisotopes as tracers.

Unit-3 Polymer chemistry (15L)[14Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Polymer chemistry]
Kinetics and mechanism of polymer processes, criteria of polymer solubility, thermodynamics of polymer solutions, polymer characterization, molecular weight of polymer (number average and weight average), methods of molecular weight determination, properties of polymers and applications.

Unit-4 Electrochemistry (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Electrochemistry]

Sign convention-American, European and IUPAC; Polarization ,demonstration of polarization, elimination of polarization ,Decomposition Potential-Factors ,applications and measurement of Decomposition potential and Overvoltage and factors affecting overvoltage, basic principle of polarography, origin of different types of current; Ilkovic equation. polarographic wave equation, Importance of E1/2 in polarography ,Dependence of E1/2 on pH and complexing agents(derivation),Pilot ion method and standard addition method to determine concentration.

Reference books: PCH2803: Physical Chemistry (Theory)

Core Reference books

(1) Statistical thermodynamics by M.C.Gupta ,revised 2nd edition, New Age International publishers
(2) Nuclear and radioactive chemistry by B K Sharma, Krishan Prakashan (2014)
(3) Advanced physical chemistry by Gurdeep Raj 35th revised edition, Goel publishing house
(4) Polymer science by Gowariker, New Age International, reprint 1986.

Other Reference books: PCH2803: Physical Chemistry (Theory)

(1) Physical chemistry by W.J.Moore, 5th edition, orient longman private ltd.
(2) Textbook of physical chemistry by S. Glasstone, D. Van Nostrand company, inc., 1946.
(3) Textbook of physical chemistry by Peter AtkinsJulio de and Paula, 9th edition, oxford press.
(4) Advanced physical chemistry by J.N.Gurtu, A.Gurtu, 11th edition , Pragati prakashan.
(5) Thermodynamics,statistical thermodynamics and kinetics by T.Engle and P.Reid, Pearson India.
(6) Statistical Thermodynamics - Fundamentals and Applications by NORMAND M. LAURENDEAU , Cambridge University Press 2005
(7) Polymer science by Gowariker, New Age International, reprint 1986.
(8) Textbook of Polymer Science By 3rd Edition edition (2 May 1984), Wiley-Blackwell.
(9) Principles of Polymer Science, 2nd edition by Bahadur&Sastry, Alpha Science.
(10) Polymer science & technology by Fried, 3rd edition, Printece-Hall.
(11) Polymer Chemistry: An Introductionby Malcolm P. Stevens, 3rd edition,Addison-Wesley *Publishing* Company.
(12) Essentials of Nuclear Chemistryby Arnikar, New Age Internation
(13) Nuclear and radio chemistry by J.W. Kannedy, G.Friedlander, 3rd edition, Wiley.
(14) Modern Electrochemistry by Bockris and Reddy Fred W. Billmeyer Jr.,

Semester: 2
Course Name: Analytical Chemistry (Theory) **Course Code: PCH-2804**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: To use different techniques like liquid-liquid extraction, counter current extraction, digestion and solid phase extraction for sample preparation

CO2: To identify as well as remember the principles and theory of chromatography and to apply them to interpret and use data generated by specialized chromatographic techniques such as GC, HPLC, HPTLC and IEC

CO3: To apply the principles of electrochemistry in the quantitative analysis of various ionic solutions using different types of ion selective electrodes

Unit-1 Sample Preparation Techniques (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Sample Preparation Techniques]

Liquid-liquid extraction/solvent extraction-partition coefficient, distribution ratio and percent extraction.Solvent extraction of metal ions-ion association complexes and metal chelates, multiple batch extraction, Craig's counter-current distribution. Accelerated and Microwave assisted extraction, protein precipitation and solid phase extraction (SPE).

Unit -2 Basic Principles of chromatography (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Basic Principles of chromatography]

Chromatographic Methods Principles of chromatography, classification of chromatographic techniques based on mechanism of retention, configuration, mobile and stationary phase. Importance and meaning of terms –Partition Ratio, Retention Time and Dead Time, Capacity Factor and Selectivity Factor, Efficiency of separation- plate theory (theoretical plate concept) and rate theory (Van Deemter equation).

Unit -3 Specialized chromatographic techniques (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Basic Principles of chromatography and Specialized chromatographic]

Principles, instrumentation and applications of Gas Chromatography, HPLC, HPTLC and Ion exchange chromatography .

Unit -4 Electro-Analytical Chemistry (15L)[14Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Electro-Analytical Chemistry]

pH measurement with glass electrode, working of glass electrode, mechanism of pH measurement, calibration of glass electrode, errors in pH measurement. Classification, principle, properties and design of ion selective membrane electrodes- Glass electrodes for ions other than H^+, Solid precipitate electrode and single crystal electrode, Liquid ion exchange electrode, ion molecular sieve electrodes. Ion, Gas-sensing probes and enzyme substrate electrodes.

Reference books: PCH2804: Analytical Chemistry (Theory)

(**Core Reference Books:**

(1) Analytical Chemistry, by Gary D. Christian, 6th Edition, John Wiley and Sons Inc. New Jersey.
(2) Quantitative Chemical Analysis, by Daniel C. Harris, 5th Edition, W.H. Freeman and Company, New York.
(3) Fundamentals of Analytical Chemistry by Crouch, West and Skoog,9th edition , Brooks/Cole (2013)

Other Reference books: PCH2804: Analytical Chemistry (Theory)

(1) A Practical handbook of preparative HPLC by Donald Wellings, Elsevier, 2006.
(2) Ion-pair chromatography: Theory and Biological and Pharmaceutical Applications (Chromatographic Science), Milton Hearn (editor), Marcel and Dekker Inc. (1985).
(3) Practical Aspects of Gas Chromatography/Mass Spectroscopy by Gordon M. Message, John Wiley & Sons, 1984.
(4) Modern Practice of Gas Chromatography by Robert L. Grob and Eugene F. Barry, 3rd edition, Wiley-Interscience, 1995.
(5) Basic Gas Chromatography by Harold M. McNair, James M. Miller, John Wiley and Sons, 2008.
(6) Analytical gas Chromatography by Walter Jennings, Eric Mittlefehldt and Philip Stremple, second edition, Elsevier Science, 1997.
(7) Modern HPLC for practicing scientists by Michael W. Dong, Wiley Interscience, 2006.

Semester: 2
Course Name: Inorganic Chemistryand Organic Chemistry (Practicals) **Course Code: PCH-2805L**
No. of Credits: 04 **Learning Hours: 2X3 Hours**

Course Outcomes:
CO1: To prepare and determine the purity of double and complex salts

CO2: To perform the colorimetric estimation of some transition metals using complexation

CO3: To apply the principles of organic spotting to analyse ternary organic mixtures

(A) Inorganic Chemistry (practicals)
II. Course Content

(1) Preparation and determination of purity of double and complex salts. At least seven preparations should be done.

(2) Colourimetric estimation of any four out of Cu, Mn, NO_2, Ni, P, Fe, V, Ti, Cr, Co.

References books: PCH2805L (A) : Inorganic Chemistry(Practicals)

(1) Vogel's Textbook of Quantitative Chemical Analysis, 6th Edition, 2002.

(2) Advanced Practical Inorganic Chemistry, Gurdeepraj, Goel Publishing House, 2001.

(3) An Advanced Course in Practical Chemistry, A.K. Nad, B. Mahapatra, A. Ghosal, New Central Book Agency, 2004

(B) Organic Chemistry (practicals)
II. Course Content
Mixture analysis: ternary mixture to be given. (S+S+S)or (L+L+L). Type determination. Separation by physical and chemical methods. (both permitted in case of liquids).

References books: PCH2805L(B): Organic Chemistry (Practicals)

(1) A text book of practical organic chemistry – A. I. Vogel

(2) Practical organic Chemistry – Mann and Saunders

(3) A handbook of quantitative and qualitative analysis – H. T. Clarke

(4) Comprehensive Practical Organic Chemistry : Qualitative Analysis V K Ahluwalia& S. Dhingra.

(5) Comprehensive Practical Organic Chemistry : Preparations and Quantitative Analysis, V KAhluwalia& R. Aggarwal Universities Press.

(6) An Advance Course in practical Chemistry, A K. Nad, B. Mahapatra and A. Ghoshal.

Semester: 2
Course Name: Physical Chemistry and Analytical Chemistry (Practicals) **Course Code: PCH-2806L**
No. of Credits: 04 **Learning Hours: 2X3 Hours**

Course Outcomes:
CO1: Use the conductometer to verify Ostwald's dilution law and Debye-Huckel-Onsager's equation

CO2: Use the potentiometer to perform acid-base titrations, precipitation titration and redox titrations

CO3: Use the pH meter to do determine the dissociation constant of a weak acid and the ionization constant of a tribasic acid

CO4: Perform experiments based on kinetics

CO5: Perform experiments based on distribution coefficient

CO6: Perform environmental analysis like determination of DO, COD

CO7: Perform titrations to determine the assay of drugs and other commercial preparations

CO8: Perform water and oil analysis

(A) Physical Chemistry (practicals)

II. Course Content

(1) **Conductometry**

(a) Test of validity of Ostwald's dilution law and determination of dissociation constant of weakelectrolyte like CH_3COOH & $ClCH_2COOH$

(b) Verification of Debye-Huckel-Onsager's equation in case of strong electrolytes like HCl, KCl, NaCl.

(2) Potentiometry

(a) Titration of dibasic acid like malonic, oxalic, succinic acid with NaOH and find the dissociation constant of acid.

(b) Precipitation titration → Titration of halids with $AgNO_3$.

(c) Redox titration Ferrous ammonium sulfate –$KMnO_4$, $K_2Cr_2O_7$.

(3) pHmetry

(a) Determination of dissociation constant of weak acid like acetic and monochloroaceticacid

(b) To titrate phosphoric acid solution against alkali and to calculate the first, second and thirdioniozation constant of it.(0.05M phosphoric acid, 10% neutral $CaCl_2$, 0.5N NaOH)

(4) Kinetics

(a) Determination of the effect of (a) Change of temperature (b) Change ofconcentration of reactant and catalyst and (c) Ionic strength of the media on the velocity constant of an acid hydrolysis of an ester.

(b) Determination of order of reaction between $K_2S_2O_8$ and KI by a fractional change method

(5) Distribution method

(a) Determination of the formula of the complex formed between cupric ion and ammonia by distribution method.

Referencesbooks: PCH2806L (A): Physical Chemistry (Practicals)

(1) Practical physical chemistry –J.B.Yadav
(2) Practicals in physical chemistry – P.S.Sindhu
(3) Experimental physical chemistry – R.C.Das, B.Behera
(4) Experiments in physical chemistry- P.H.Parsania, F. Karia

(B): Analytical Chemistry (practicals)

II. Course Content

(1) Determination of saponification value of oil.
(2) Determination of iodine value of oil.
(3) Determination of iron by chloride extraction by solvent extraction process.
(4) Determination of dissolved oxygen.
(5) Determination of chemical oxygen demand.
(6) Determination of iron in iron tablets.
(7) Simultaneous estimation of chromium (III) and iron (III) by EDTA titration.
(8) Simultaneous estimation of calcium (II) and zinc (II) by EDTA titration.
(9) Simultaneous estimation of lead (II) and magnesium (II) by EDTA titration.
(10) Separation of aminoacids/ dyes/ drugs by TLC.
(11) Determination of cation content in hard water by ion exchange chromatography.

References books: PCH2806L (B): Analytical Chemistry (Practicals)

(1) Analytical Chemistry Practice, John H. Kennedy, Saunders College Publishing, Second Edition 1990.
(2) Vogel's Textbook of Quantitative Chemical Anal

Semester: 3
Course Name: Natural Products and Biomolecules (Theory) **Course Code: PCH-3801**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Remember the classification, synthesis, biosynthesis and biochemical roles of some natural pigments, alkaloids, vitamins, steroids, hormones, terpenoids and carotenoids

CO2: Employ the fundamentals of organic reaction mechanisms for the structure elucidation of some natural pigments, alkaloids, vitamins, steroids, hormones, terpenoids and carotenoids

Unit 1 Natural pigment **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Natural pigment]

Natural colouring matter, general classification, method of synthesis, biosynthesis studies of anthocyanins (cyanine) flavones (chryosin) and flavanol (Querecetin)
Porphyrin-structure, spectral properties and synthesis, general and structure determination of Haemoglobin, chlorophyll and Bilirubin.

Unit 2 Alkaloids and vitamins **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Alkaloids and vitamins]

Alkaloids: General biogenetic studies of alkaloids, chemistry of quinine, morphine, reserpine and colchicine
Vitamins : Introduction, Classification, synthesis and biochemical function of vitamin B(Thiamine), Vitamin H and α -tocopherol (Vitamin E), vitamin C.

Unit 3 Steroids and hormones **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Steroids and hormones]

(A) Brief account on the chemistry of steroids: structure of cholesterol and ergosterol, Bile acid-Lithiocholic (No synthesis).
(B) Chemistry of androgens (Androsterone and testosterons), oestrogens-Oestrone and gestrogens-Progesteron, their synthesis and biochemical role

Unit 4 Terpenoids and carotenoids **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Terpenoids and carotenoids]

(A) Introduction to Terpenes, Classification, nomenclature, general methods of structure determination, chemistry and synthesis, Monoterpenes-Geraniol & terpineol, Sesquiterpenes-farnesol, zingeberine Diterpenoids- abietic acid and gibberellic acid

(B) Carotenoids : Introduction, classification of carotenoids and B-carotenoids –(Structural elucidation and synthesis of β-carotene – their uses.)

Reference books: PCH 3801: Natural Products and Biomolecules (Theory)

(1) Organic chemistry vol I & II (sixth edition) I.L.Finar
(2) Chemistry of vitamins-S.F.Dyke
(3) Chemistry of natural products by Bantely, Vol 1-10
(4) L.J.Wade Jr. Organic chemistry, Prentice Hall, England cliffs, 1987
(5) Chemistry of Natural products vol I & II by O.P.Agrawal
(6) Essentials of medicinal chemistry , eds., Korolkovas and Burkhaltar, J.H., John wiley & sons .
(7) Text book of Organic medicinal and pharmaceutical chemistry by Wilson and Gisvold.
(8) Synthetic drugs by O. D. Tyagi.

Semester: 3
Course Name: Medicinal chemistry (Theory) **Course Code: PCH-3802**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Identify the classes, synthetic routes and mechanism of action of drugs that act as antibiotics, psychoactive agents, antimalarials, antituberculer agents, cardiovascular, diuretics and hypoglycemic agents

CO2: Remember the structure activity relationship (SAR) of some drugs that act as antibiotics, psychoactive agents, antimalarial and antituberculosis agents, cardiovascular, diuretics and hypoglycemic agents and co-relate their chemical structure to the biological activity

Unit 1 Antibiotics (15L)[14 Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Antibiotics]

(A) The β-lactum antibiotics: Penicillin and cephalosporin
Penicillin: Structure determination, Mode of action and Structure activity relationship (SAR), Synthesis of Penicillin V.
Cephalosporin: First, second, third and fourth generation of cephalosporin's their SAR, mode of action, Synthesis of Cephalosporin

(B) Non lactum antibiotics:Tetracycline, Chloroamphenicol
Tetracycline: General characteristics, Mode of action and SAR of tetracyclines
Chloroamphenicol: Structure, SAR, Mode of action and its synthesis.
(C) Non classifiable antibiotics (only structure and therapeutic uses)
Bacitracin, vancomycin, nalidixic acid, norfloxacin and ciprofloxacine
(Only name and structures)

Unit 2 Psychoactive drugs (15L)[14 Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Psychoactive drugs]

Introduction, classification, structure activity relationship (SAR), Mode of action of
(A) General and local Anesthetics
(B) Sedative and hypnotics Antipsychotic drugs
(C) Antidepressant
(D) Neuroleptics and selected synthesis of the following: Thiopental, amobarbital, diazepam, zaleplon (Sonata), alprazolam, glutethimide, nikethamide, Haloperidol, Aripiprazole, procaine, lidocaine, dibucaine, Fluoxetineand escitalopram

Unit 3 Antimalarials and Antituberculosis drugs (15L)[14 Marks]
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Antimalarial and Antituberculosis drugs]

(A) Antimalarials: Modern chemotherapy of malaria, 4-amino and 8-amino quinolins, 9-amino acridine. Synthesis of mefloquines, chloroquine, primaquine and daraprim
Mode of action of antimalarial agents SAR of antimalarial agents.

(B) Anti tuberculosis: Introduction, classification and mode of action, drug resistance tuberculosis. Synthesis of only the following drugs: Isoniazid (INH), Ethionamide, Ethambutol, DDS (Dapsone).

Unit 4 Cardiovascular, diuretics and hypoglycemic agents (15L)[14 Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Cardiovascular, diuretics and hypoglycemic agents]

Cardiovascular agents: Introduction, classification, therapeutic drug categories and mode of action of any three category.

Diuretics: Introduction, types of diuretics, mode of action (three categories)

Hypoglycemic agents: Introduction, classification of Type 2 hypoglycemic agents (sensitizers)

Synthesis of amyl nitrate, diltiazim, atenolol, methyl dopa, tolbutamide, chloropropamide, glibenclamide, acetazolamide, chlorothiazide, furosemic and ethacrycnic acid.

Reference books: PCH 3802: Medicinal Chemistry (Theory)

(1) Burger's medicinal chemistry and drug design (5/e) 1997, vol 1 to 5 edited by Manfred E.Woltt (John wiley and sons Mc. New york)
(2) Principles of medicinal chemistry by William A. Foye (ied), lea and febiys (Philadelphia)
(3) Principles of medicinal chemistry vol I & II (5/e) F.S.kadam, K.R. Mahadic and K.G.Bohra (Nirali publication)
(4) Medicinal chemistry by Ashutosh kar
(5) The organic chemistry of drug synthesis vol I, II and III (1980) ed by D. lednicer and L.A. mitscher (Johyn wiley and sons, New york)
(6) Wilson and Gisvold text book of organic medicinal and pharmaceutical chemistry (5/e,1982) by Robert Doerge (J.B. lippincoff company, phaladophia/ Toppan co.Ltd, Tokyo)
(7) Topics in medicinal chemistry vol I & II by rabinowitz Myerson (interscience 1968)
(8) The pharmaceutical basis of therapeutics by Geoman and Gilman (Mcmillan co.)

St. Xavier's College (Autonomous)
Ahmedabad-09
M.Sc. Organic Chemistry

Semester: 3
Course Name: Organic Spectroscopy (Theory) **Course Code: PCH-3803**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Remember the principle, instrumentation of 13C-NMR and 1H-NMR as well as their advanced applications in DEPT, 2D-NMR. NOE, COSY and HETCOR

CO2: Identify the various methods for the ionization of molecules, recognize the different modes of fragmentation, the rules for molecular mass determination as well as the instrumentation of a mass spectrometer

CO3: Employ the knowledge of spectroscopy and spectral interpretations to determine the structure of molecules

Unit 1 ^{13}C-NMR **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to ^{13}C- NMR]

^{13}C NMR Spectroscopy: Introduction, Instrumentation and instrumental problems, ^{13}C-chemical shifts, factors affecting chemical shift (aliphatic, olefinic, alkyne, aromatic, heteroaromatic and carbonyl carbon), coupling constants, proton coupled ^{13}C - spectra, proton decoupled ^{13}C- spectra. Off- resonance decoupling, DEPT technique.

Unit 2 Mass Spectroscopy **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to and Mass spectrometry]

Mass Spectroscopy: Introduction, Instrumentation, Ionization technique-EI, CI, FD and FAB, factors affecting fragmentation, ion analysis, ion abundance, mass spectral fragmentation of organic compounds, common functional groups, Detection of molecular ion peak, base peak, metastable ion peak, McLafferty rearrangement, nitrogen rule, high resolution mass spectrometry, examples of mass spectral fragmentation of organic compounds with respect to their structure determination. Hyphenated mass spectroscopy.

Unit 3 NMR **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to NMR]

FT NMR instrumentation and principle, mechanism of spin-spin coupling, Complex spectra and spin-spin splitting, second order effects on the spectrum, A2, AB, AX, AB2….AX2, A2B2….A2X2, AMX, ABX spin systems, simplification of second order spectra.
Nuclear overhauser effect, 2D NMR (COSY and HETCOR) applications, HSQC, HMQC

Unit 4 Structural elucidation **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to UV, IR, PMR, CMR and mass spectroscopy]

Structural elucidation of drug molecules based on joint application of UV, IR, PMR, CMR and mass spectroscopy.

Reference books: PCH 3803 : Organic Spectroscopy (Theory)

(1) Spectroscopic methods in organic chemistry, D.H.Williams and Tan fleming.

(2) Spectrometric identification of organic compounds, T.C.Morril R.M.Silverstein and G.Bassler, 6th edition, John Wiley and sons

(3) Introduction to spectroscopy, D.L.Pavia, G.M.Lampman and G.S.Kriz, 3rd edn, Harcourt college publishers.

(4) Organic spectroscopy by W.Kemp

(5) Organic spectroscopy by P.S.Kalsi

Semester: 3
Course Name: Industrial Chemistry (Theory) **Course Code: PCH-3804**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Describe unit processes such as nitration, sulphonation etc as well as unit operations such as crystallization, filtration etc

CO2: remember the 12 principles of green chemistry, recognize green reagents, solvents, catalysts and use them efficiently for designing green processes

CO3: Give an outline regarding the raw materials used in the cosmetic industry especially in baby care, hair care, dental care and synthetic perfumes

CO4: Describe the classification of textile and non-textile dyes; their synthesis and the basic operations in the process of dyeing

Unit 1 Unit Processes and Operations **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental terms related toUnit Processes and Operations]

(A) Nitration, oxidation, sulphonation, reduction, halogenations,

(B) Filtration, extraction, crystallization, drying, distillation and industrial chemicals derived from benzene, naphthalene and anthracene using unit process.

Unit 2 Green chemistry **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Green chemistry]

Green chemistry -12 principles of green chemistry
Green solvents- aqueous phase reactions, Wurtz reaction, witting-Horner reaction, Michael reaction
Solid phase reactions: halogenation, aldol condensation, grignard reaction.
Ionic liquid as green solvent: hydrogenation, diels-alder reaction, o-alkylation and N-alkylation.
Introduction: green catalysts, green reagents, Microwave Madelung Indole synthesis, Enzymatic Conversion of sucrose to ethanol.

Unit 3 Cosmetics

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to cosmetics, Perfumes]

(A) Classification, Raw materials for cosmetics, Manufacturing of various cosmetic products Baby Care Product , Dental Products, Hair Care product

(B) Synthetic Perfumes Definition, classification, synthesis and uses of : ester of cinnamic acid, linalool, phenyl ethyl alcohol, civet one, musk ambrette, alpha and β-ionones, alpha and β-irones.

Unit 4 Synthetic Dyes and Pigments (15L)[14 Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Synthetic Dyes and Pigments]

Introduction, Relation between colour and chemical constitution of dye, hot and cold brand dyes, types of fibers to be dyed, Basic operations in dyeing, Methods of Dyeing of direct, mordant and vat dyes, Classification of dyes according to their applications and chemical constitutions, Fluorescent brightening agents. synthesis of Rosaniline, Tinopal BV, Rhodamine 6G (Red).

Non textile dyes: Leather dyes, Paper dyes, Food colours, Cosmetic dyes, Medicinal dyes.

Pigments: Introduction, Inorganic pigments, application of titanium pigments, organic pigments, chemistry of copper pthalocyanine.

Reference books: PCH 3804: Industrial Chemistry (Theory)

(1) Unit processes in organic synthesis by P.H.Groggins

(2) Industrial Chemical process by R.N.Shreve

(3) Riegels handbook of industrial chemistry ed by James and Kent

(4) Dryden's outlines of chemical Technology M.Gopal Rao

(5) The chemistry of oils & fats,F.D.Gunstone,Blackwell Pub.

(6) Baileys Industrial oils & fats products,Vol 1-5,John Wiley & Sons

(7) Essential Oils,Vol 1-7,D.Gunther,R.E.Krigger Pub Comp.,New York

(8) Cosmetic Science & Technology,Vol 1 &2,Wiley Interscience,New York

(9) Cosmetics,Soaps& Perfumes,W.A.Poucher,Chapman Hall,London & New York

(10) Industrial Chemistry by B K Sharma, Sixteenth edition,GOEL publishing house,Meerut.

(11) Hand book of Synthetic dyes and Pigments, Vol I, II , III, by K M Shah, second edition, Multi-tech publishing co. Mumbai.

Semester: 3
Course Name: Organic Preparations (Practicals) **Course Code: PCH-3805L**
No. of Credits: 04 **Learning Hours: 2X3ours**

Course Outcomes:
CO1: Prepare compounds of industrial importance by using name reactions such as Sandmeyer reaction, Pechmann reaction, Skraup synthesis, Riemer-Tiemann reaction, Kolbe-smith reaction, Claisen-smith synthesis, Hoffman reaction, Diels-alder reaction, Green –bromination

CO2: Use the fundamentals of organic reaction mechanism to determine the mechanisms of these name reactions

CO3: Characterize the compound prepared

Organic Preparations (Practicals)
II. Course Content
Preparation of industrially important compounds by following name reactions (mechanism, purification and characterization of the synthesized compounds)

1. Sandmeyer reaction
2. Pechmann reaction
3. Skraup synthesis
4. Riemer-Tiemann reaction
5. Kolbe-smith reaction
6. Claisen-smith synthesis
7. Hoffman reaction
8. Diels-alder reaction
9. Green –bromination

References books: PCH 3805L : Organic Preparations (Practicals)

(1) A text book of practical organic chemistry by A. I. Vogel
(2) Practical organic Chemistry by Mann and Saunders
(3) A handbook of quantitative and qualitative analysis by H. T. Clarke
(4) Comprehensive Practical Organic Chemistry : Qualitative Analysis by V K Ahluwalia & S. Dhingra.
(5) Comprehensive Practical Organic Chemistry : Preparations and Quantitative Analysis by VK Ahluwalia& R. Aggarwal Universities Press.
(6) An Advance Course in practical Chemistry, by A K. Nad, B. Mahapatraand A. Ghoshal.

Semester: 3
Course Name: Organic Preparations (Practicals) **Course Code: PCH-3806L**
No. of Credits: 04 **Learning Hours: 2X3 Hours**

Course Outcomes:
CO1: Use the fundamental analytical techniques to determine the assay of drugs

CO2: Perform non-aqueous titrations

CO3: Determine the nitrite value of the given solution

CO4: Extract casein from milk

Organic Estimations (Practicals)
II. Course Content

Estimations and Extractions

(1) Drug assay (estimation of sulpha drug)

(2) Non-aqueous titration (Nicotine estimation in $HClO_4$)

(3) Nitrite value

(4) Extraction of Casein from Milk

References books: PCH 3806L: Organic Estimation (Practicals)

(1) Quantitative analysis by Arther I.Vogel

(2) Quantitative analysis by V.K.Ahluwalia

(3) Quantitative analysis by Mann and sanders

Semester: 4
Course Name: Advnaced Organic Chemistry (Theory) **Course Code: PCH-4801**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Illustrate the applications of the Woodward-Hofmann, FMO and PMO method in determination of the stereochemical outcome of pericyclic reactions such as electrocyclic, cycloaddition and sigmatropic reactions

CO2: Use the fundamentals of conformational analysis to determine the stable conformations of simple, substituted and heterocyclic saturated ring systems

CO3: Describe the mechanism of aliphatic electrophilic substitution reactions and the properties of the electrophiles involved

CO4: Illustrate the modern types of organic synthetic methodologies and also describe some modern name reactions

Unit 1 Pericyclic reactions **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Pericyclic reactions]

Introduction, classification of pericyclic reactions, stereochemistry, molecular orbital symmetry, frontier orbitals of ethelene, 1,3 –butadiene, 1,3,5-hexatriene and allyl system, F.M.O. and PMO approach to cycloaddition and electrocyclic reactions: Generalization of wood-ward Hoffmann rule, sigmatropic rearrangement-suprafacial and antrafacial shifts of H.Stereoselectivity in sigmatropic rearrangement, enantioselectinity in pericyclic reactions.

Unit 2 Conformational analysis **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Conformational analysis*]*

Conformational descriptors : Factors affecting conformational stability of molecules, conformational analysis of mono and disubstituted cyclohexane, other cycloalkanes, conformations in six-membered rings containing heteroatoms
Polycyclic compounds decalins, perhydro anthracene, perhydro phenanthrene, norbornane
Bridged systems and strained carbocycles

Unit 3 Aliphatic Electrophilic substitutions **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Aliphatic Electrophilic substitutions]

Aliphatic Electrophilic substitutions: SE1 SE2 and SEi Bimolecular mechanisms, electrophilic substitution accompanied by double bond shifts, migration of double bonds, Hydrogen, Halogen, nitrogen, carbon sulphur and metal as electrophiles and their reactions (selective reactions only). Metal and Halogen as leaving group.

Unit 4 Modern Synthetic Methods (15L)[14 Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Modern Synthetic Methods]

Baylis-Hillman reaction, Henry reaction, Nef reaction, Kulinkovich reaction, Ritter reaction, Sakurai reaction, Tishchenko reaction, Ugi reaction, Noyori reaction. Brook rearrangement. Tebbe olefination.

Introduction to chemistry of multicomponent reactions and Click reactions. Domino/cascade reactions: Introduction with one example.

Reference books: PCH 4801 : Advanced Organic Chemistry (Theory)

(1) Advanced organic chemistry - Reaction mechanism and structure, Jerry March, John Wiley publication.

(2) Structure and mechanism in organic chemistry, C. K. Ingold, Cornell UniversityPress.

(3) Advance organic chemistry by Francis A. Carey

(4) Named organic reactions by Laslo Kurti

(5) Advanced organic chemistry, Part I & II F. A. Carey and R. J. Sundberg, Plenum.

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Semester: 4
Course Name: Advnaced Organic Synthesis (Theory) **Course Code: PCH-4802**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Describe the principles involved in protection of functional groups during organic synthesis

CO2: Remember the fundamentals of the disconnection approach and identify the principles of retrosynthesis to carry out C-C disconnections

CO3: Use basic organic chemistry to remember the synthesis of saturated and aromatic heterocyclic rings

Unit 1 Protection of groups **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Protection of groups]

Principle of protection of hydroxyl, amino, carbonyl, carboxylic acid with different reagents and their deprotection, synthetic equivalent groups, synthetic analysis and planning, control of stereochemistry.

Unit 2 Disconnection approach **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Disconnection approach]

An introduction to synthesis, and synthetic equivalents, disconnection approach, functional group inter-conversions, the importance of the order of events in organic synthesis one group C-X and two group C-X disconnections, chemo-selectivity, reversal and polarity.

Unit 3 One group C-C disconnections **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to : One group C-C disconnections]

Alcohols and carbonyl compounds, region-selectivity, alkene synthesis, use of acetylenes and aliphatic nitro compounds in organic synthesis.

Unit 4 Ring synthesis of heterocyclic compounds **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to ring synthesis]

Saturated heterocycles, synthesis of 3, 4, 5, and 6-membered rings, aromatic heterocycles in organic synthesis.

Reference books: PCH 4802 : Advanced Organic Synthesis (Theory)

(1) Organic synthesis : the disconnection approach by stuart Warren (wiley student edition)

(2) Organic chemistry- Clayden, Greeves, Warren and Wothers, (oxford press)

(3) Structure and mechanism in organic chemistry, C. K. Ingold, Cornell UniversityPress.

(4) Advanced organic chemistry, Part I & II F. A. Carey and R. J. Sundberg, Plenum.

Semester: 4
Course Name: Biorganic Chemistry (Theory) **Course Code: PCH-4803**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Remember the fundamentals of biochemistry in the context of amino acids, peptides and nucleic acids

CO2: Summarise the basic biochemical reactions and processes involved in metabolism of carbohydrates, proteins and fatty acids

CO3: Explain the structure, functions and complex reactions/processes undergone by nucleic acids, carbohydrates and vitamins

Unit 1 Fundamentals of Biochemistry **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Fundamentals of Biochemistry]

Introduction of Biochemistry, Amino acids: peptides, primary, secondary, tertiary, and quaternary structure of proteins. Nucleic acids: Base pairing, double helices, DNA replication, transcription andtranslation, Enzymatic hydrolysis of proteins to peptides; Amino acid sequencing; amino acidmetabolism (biosynthesis and degradation).

Unit 2 Metabolism and Metabolic Reaction **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Metabolism and Metabolic Reaction]

Overview and important relationships between-glycolysis, Bioenergetics And Bio energetic principles, oxidative phosphorylation process, ATP synthetize, photo phosphorylation.
Fatty acid metabolism: Biological importance of fatty acids and lipids, even chain and odd chain fatty acids, saturated and unsaturated fats, ketone bodies, fatty acid metabolism, calorific value of foods, biological membranes, properties and function of lipid bilayers and liposomes.

Protein-related transformations: urea cycle, uric acid and ammonia formation.

Unit 3 Nucleic Acids **(15L)[14 Marks]**
[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Nucleic Acids]

Chemical and enzymatic hydrolysis of nucleic acids; Structure and function of mRNA, tRNA, rRNA; Polymorphic nature of DNA, B-and Z-DNA, multi-stranded DNA; DNA sequence determination by chemical and enzymatic methods, Genetic code–origin, salient features, Geneexpression transcription and translation; Gene mutation and carcinogenesis

Unit 4 Selected topics in Carbohydrate and Vitamin Chemistry **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to selected topics *in Carbohydrates and Vitamins]*

Carbohydrates: classification and stereochemistry, biologically important hexose derivatives, nomenclature of disaccharides, structure and role of some homo and hetero polysaccharides, polysaccharides: starch and cellulose, glycosides: salicin, indican and amygldalin, glucoconjugates : proteoglycans, glycoproteins and glycolipids.
Vitamins : Synthesis and biochemical function of vitamin A, vitamin D and vitamin K.

Reference books: PCH 4803:Bioorganic Chemistry(Theory)

(1) Albert L. Lehninger, David L. Nelson, Michael M. Cox., Principles of Biochemistry, CBS Publishersand Distributors, 1993.

(2) Lubert Stryer, Biochemistry, W. H. Freeman and Company, 4th edition, 1995.

(3) Christopher K. Mathews and K. E. Von Holder, Biochemistry, Benjamin/Cummings, 1990.

(4) Eric E. Conn, Paul K. Stumpf, George Breening and Roy H. Doi, Outlines of Biochemistry, 5th edition,John Wiley and Sons,1987.

(5) Organic Chemistry by F. A. Carey and R. J. Sundberg, (Eds) 3rd Edition, Part B. Plenum/Rosetta,1990.

(6) I.Fleming, Selected Organic Synthesis, John Wiley and sons, 1982.

(7) Atta-ur-Rehman, Studies in Natural Products Chemistry, Vol.1 and 2, Elsevier, 1988.

(8) T.Lindberg, Strategies and Tactics in Organic Synthesis, Academic Press, 1984.

(9) E. J. Corey and X-M. Cheng, Logic of Chemical Synthesis, John Wiley, 1989.

(10) H. Pape and J. H. Rehm, (eds): Biotechnology, A Comprehensive Treatise,Vol. 1-8, VCH, 198

(11) Principles of biochemistry –Donald J.Voet, Judish G. Voet, Charlotte w. Pratt (John Willey and sons)

(12) Lehninger principles of biochemistry- David L.Nelson and Michael M. (Palgrave Macmillan/ w.h. freeman company new york)

(13) Biochemistry – U.Satyanarayana Baro and allied P.Ltd., Kolkata

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Semester: 4
Course Name: Selected topics in medicinal Chemistry (Theory) **Course Code: PCH-4804**
No. of Credits: 04 **Learning Hours: 60 Hours**

Course Outcomes:
CO1: Recognize the principles involved in drug design and describe the different aspects of medicinal chemistry such as pharmacokinetics, pharmacodynamics, drug activity theories, clinical trials etc

CO2: Describe the quality control and assurance methods used in the pharma industries as well as the different routes of administration and dosage forms

CO3: Identify the use of computers in chemistry and locate the databases available for literature survey in medicinal chemistry

CO4: Compile the nanomaterials and nanotechnology-based devices used for tackling problems in pharmacokinetics and pharmacodynamics

Unit 1 Drug design **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Drug design]

Introduction, naming of organic medicinal compounds, literature of medicinal chemistry, development of new drugs, procedure followed in drug design, concept of lead compound and lead modification, pro drugs, soft drugs, phase I, II and III clinical trials, structure activity relationship, theories of drug activity : occupational theory, rate theory, induced fit theory, quantitative structure activity relationship, history and development of QSAR. Concept of drug receptors, elementary treatment of drug receptor interactions, physio chemical parameters lipophilicity, partition coefficient, electronic ionization constant, concept of 3-D QSAR.

Unit 2 Pharmacokinetic and pharmacodynamics **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Pharmacokinetic and pharmacodynamics]

Pharmacokinetics : introduction to drug absorption, distribution, metabolism, elimination. important pharmacokinetic parameters in defining drug deposition and in therapeutics, uses of pharmaceutics in drug development process
Pharmacodynamics: Introduction, elementary treatment of enzyme stimulation, enzyme inhibition, drug metabolism, biotransformation, significance of drug metabolism in medicinal chemistry.

Unit 3 Quality control and computer applications in pharm industries **(15L)[14 Marks]**

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Quality control and computer applications in pharm industries]

(A) Dosage forms, Quality control and application of computers in chemistry

Dosage forms, types of dosages, different roots of administration, quality control of drugs pharmacopias, modern methods of pharmaceutical analysis.

(B) Computer in chemistry

Use of computer in chemistry and industry
Important websites for data search chemistry
Information about online journals for chemistry

Unit 4 Nano science and Medicinal chemistry (15L)[14 Marks]

[Prerequisites or topics for Self Study: Basic terms and fundamental aspects related to Nano science and Medicinal chemistry]

Overview, Medicinal use of nanomaterials-Drug delivery
Protein and peptide delivery –cancer, surgery, visualization, nanoparticle targeting
Medical application of molecular nanotechnology-nanorobots, cell repair machines, nanonephrology.

Reference books: PCH 4804 : Selected topics in Medicinal Chemistry (Theory)

(1) Burger's Medicinal Chemistry and Drug Discovery (5/e), 1997, Vol. 1, 2, 3, 4,5, Edited by ManFred E. Wolff (John Wiley & Sons, inc., New York).

(2) Wilson and Gisvold's Text-book of Organic Medicinal and Pharmaceutical Chemistry (5/e, 1982) by Robert F. Doerge (J. B. Lippincott Company, Philadelphia/Toppan Co. Ltd., Tokyo).

(3) Principles of Medicinal Chemistry, Vol. I & II (5/e), by S. S. Kadam, K. R. Mahadik, K. G. Bothra (Nirali Prakashan).

(4) QSAR: quantitative structure-activity relationships in drug design by Jean-Luc Fauchère. ISBN:084515141X, 9780845151419.

(5) QSAR : Hansch analysis and related approaches By Hugo Kubinyi

Semester: 4
Course Name: Industrial Training (Practicals) **Course Code: PCH-4805L**
No. of Credits: 04 **Learning Hours: 2X3 Hours**

Course Outcomes:
CO1: Apply the theories and concepts learned in real-life world

CO2: Communicate and present the applications, training and experience from the industrial training

CO3: Analyse the observations from the industrial training

Course Content

Industrial training for 21 days, report preparation, submission and presentation / viva

References: PCH 4805L: Industrial training (Practicals)

(1) Industrial Chemistry, by B. K. Sharma ,GOEL Publishing House, Meerut.

(2) Elementary Practical Organic Chemistry, Part I, II and III by Arthur I. Vogel.

(3) Practical organic Chemistry, by F. G. Mann and B. C. Saunders, 4th Edition.

(4) Instrumental Analysis, by William Kemp, 3rd Edition.

(5) The reference material provide by the industries and the books related to the area of the training.

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Semester: 4
Course Name: Organic Preparations and Estimation (Practicals) **Course Code: PCH-4806L**
No. of Credits: 04 **Learning Hours: 2X3 Hours**

Course Outcomes:
CO1: Prepare compounds of industrial importance by using starting materials having specific functional groups

CO2: Use the fundamentals of organic reaction mechanism to determine the mechanisms of these name reactions

CO3: Characterize the compound prepared

CO4: Use the fundamental analytical techniques to determine the assay of drugs

(A) Organic Preparations
(1) Preparation of cinnamic acid form benzaldehyde.
(2) Preparation of benzophenone oxime from benzophenone
(3) Preparation of anthraquinone form anthracene
(4) Preparation of 4-phenyl-6-methyl-5-carbethoxy-2-pyrimidone from urea, benzaldehyde, EAA (Ethyl Aceto Acetate)
(5) Preparation of 2-phenylindole from acetophenone and phenylhydrazine.
(6) Preparation of 2-methylbenzimidazole from *o*-phenylene diamine
(7) Preparation of 1, 1-bis-2-naphthol from 2-Naphthol.

(B) Organic Estimation

(1) Estimation of Aspirine
(2) Estimation of Isoniazid (INH)
(3) Estimation of Ibuprofen

References books: PCH 4806L: Organic Preparations and Estimation (Practicals)
(1) A text book of practical organic chemistry by A. I. Vogel
(2) Practical organic Chemistry – Mann and Saunders
(3) A handbook of quantitative and qualitative analysis by H. T. Clarke
(4) Comprehensive Practical Organic Chemistry : Qualitative Analysis V K Ahluwalia and S. Dhingra.
(5) Comprehensive Practical Organic Chemistry : Preparations and Quantitative Analysis V

K Ahluwaliaand R. Aggarwal Universities Press.
(6) An Advance Course in practical Chemistry, A K. Nad, B. Mahapatraand A. Ghoshal.
(7) Quantitative analysis by Arther I.Vogel
(8) Quantitative analysis by V.K.Ahluwalia
(9) Quantitative analysis by Mann and sanders